

Home learning activities

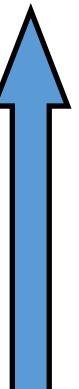
Subject
Science
Year Group
Year 10
Unit of work / Knowledge organiser
Chemical Changes - 1
Activities
<ul style="list-style-type: none"> • Complete the weekly 'Knowledge Check' through 'GCSEPod'. • Watch all 'GCSEPod' clips on the 'Chemical Changes' Unit. • Complete the 'GCSEPod' Questions assigned for this Unit of work and any additional assignments which have been set by your teacher. • Follow the 'Revision Plans' for Biology and Physics • Complete the assigned activities for the given week on the Biology and Physics revision plans
Where do you complete the work?
Use computer/phone for 'GCSEPod' or 'Seneca' and study materials.
What to do if you finish the work? (Extension activity)
<ul style="list-style-type: none"> • Sign up for 'Seneca Learning' using the 'Sign Up Guide' sheet and the special passcode: j5v9tvzq48. Complete the assignments which have been set.
These websites might help:
<ul style="list-style-type: none"> • BBC Bitesize -> Secondary -> GCSE -> Combined Science -> AQA Trilogy -> Chemistry -> Chemical Changes • www.freesciencelessons.co.uk -> GCSE Videos -> Chemistry Paper 1 -> Chemical Changes
If you are struggling with your work or if you have finished.
<p>Please email your classroom teacher directly using the email list found in the Home Learning section of the website.</p>

Year 10: Chemical Changes & Energy Changes

Section 1: Key Terms

1 Metal oxide	Metals react with oxides to produce metal oxides. This is an oxidation reaction.
2 Displacement reaction	A more reactive metal can displace a less reactive metal from a compound .
3 Oxidation	Two definitions: Chemicals are oxidised if they gain oxygen in a reaction. Chemicals are oxidised if they lose electrons in a reaction. (HT)
4 Reduction	Two definitions: Chemicals are reduced if they lose oxygen in a reaction. Chemicals are reduced if they gain electrons in a reaction. (HT)
5 Acid	A chemical that dissolves in water to produce H^+ ions.
6 Base	A chemical that reacts with acids and neutralise them. E.g. metal oxides, metal hydroxides, metal carbonate
7 Alkali	A base that dissolves in water . It produces OH^- ions in solution.
8 Neutralisation	When a neutral solution is formed from reacting an acid and alkali . General equation: $H^+ + OH^- \rightarrow H_2O$
9 pH	A scale to measure acidity/ alkalinity . A decrease of one pH unit causes a 10x increase in H^+ ions . (HT)
10 Strong acid (HT)	A strong acid is completely ionised in solution. E.g. hydrochloric, nitric and sulfuric acids .
11 Weak acid (HT)	A weak acid is only partially ionised in solution. E.g. ethanoic, citric and carbonic acids .

Section 2: Reactivity

Element	Reaction	Reactivity
12 Potassium	When potassium is added to water , the metal melts and floats. It moves around very quickly. The metal is also set on fire , with sparks and a lilac flame .	
13 Sodium	When sodium is added to water , it melts to form a ball that moves around on the surface. It fizzes rapidly .	
14 Lithium	When lithium is added to water , it floats. It fizzes steadily and becomes smaller.	
15 Calcium	Fizzes quickly with dilute acid .	
16 Magnesium	Fizzes quickly with dilute acid .	
17 (Carbon)		
18 Zinc	Bubbles slowly with dilute acid .	
19 Iron	Very slow reaction with dilute acid .	
20 (Hydrogen)		
21 Copper	No reaction with dilute acid .	

Acids produce hydrogen ions (H^+) in aqueous solutions. Aqueous solutions of alkalis contain hydroxide ions (OH^-).

Section 4: Extracting Metals

22 Very unreactive metals	Found naturally in the ground. Don't need extracting .
23 Metals less reactive than carbon	Extracted by reduction with carbon .
24 Metals more reactive than carbon	Extracted by electrolysis .

Section 5: Reactions of Acids

25 With metal	$Acid + Metal \rightarrow Salt + Hydrogen$
26 With alkali	$Acid + Metal\, Hydroxide \rightarrow Salt + Water$ (Neutralisation reaction)
27 With metal oxide	$Acid + Metal\, Oxide \rightarrow Salt + Water$ (Neutralisation reaction)
28 With carbonate	$Acid + Metal\, Carbonate \rightarrow Salt + Water + Carbon\, Dioxide$ (Neutralisation reaction)

Section 6: Making a Soluble Salt

29	Add solid metal, metal carbonate, metal oxide or metal hydroxide to an acid .
30	Add solid until no more reacts .
31	Filter off excess solid.
32	Evaporate to remove some of the water.
33	Leave to crystallise .
34	Remove all water in a desiccator/ oven .

Higher only:

A strong acid is completely ionised in aqueous solution. Examples of strong acids are hydrochloric, nitric and sulphuric acids.

A weak acid is only partially ionised in aqueous solution. Examples of weak acids are ethanoic, citric and carbonic acids.

For a given concentration of aqueous solutions, the stronger an acid, the lower the pH.

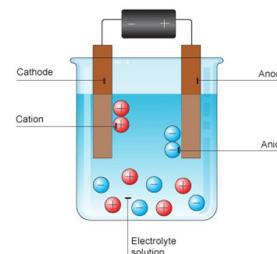
As the pH decreases by one unit, the hydrogen ion concentration of the solution increases by a factor of 10.

pH scale



Test solutions using an indicator solution or a pH probe

Year 10: Chemical Changes & Energy Changes



**Positive
Anode
Negative
Is
Cathode**

Section 7 Electrolysis key terms

38 Electrolysis	The process of splitting an ionic compound by passing electricity through it.
39 Electrolyte	An ionic compound that is molten (melted) or dissolved in water . The ions are free to move .
40 Electrode	An electrical conductor that is placed in the electrolyte and connected to the power supply .
41 Cathode	The electrode attached to the negative terminal of the power supply .
42 Anode	The electrode attached to the positive terminal of the power supply .

Section 8: What is discharged in electrolysis?

Electrolyte	Cathode	Anode
43 Molten Compound	Metal	Non-metal
44 Dissolved compound (aqueous solution)	The metal if the metal is less reactive than hydrogen . Hydrogen is produced if the metal is more reactive than hydrogen .	Oxygen is produced unless the solution contains halide ions (chloride, bromide, iodide) when the halogen (chlorine, bromine, iodine) is produced.

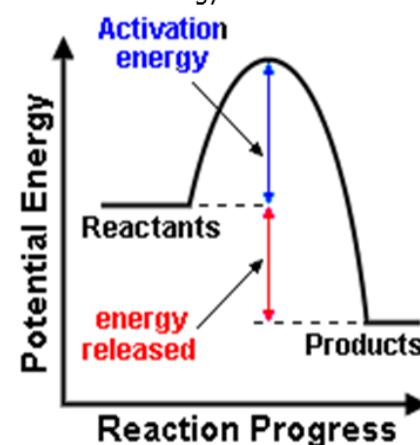
Section 9: Aluminium Electrolysis

45 Cryolite	Aluminium oxide is dissolved in cryolite to lower its melting point. This saves money on energy costs .
46 Cathode	Positive Al^{3+} ions move to the cathode . Aluminium is produced. $\text{Al}^{3+} + 3\text{e}^- \rightarrow \text{Al}$
47 Anode	Negative O^{2-} ions move to the anode . Oxygen is made. $2\text{O}^{2-} \rightarrow \text{O}_2 + 4\text{e}^-$ Wears away as the carbon anode reacts with oxygen to form carbon dioxide.

Section 7 Energy Changes Key Terms

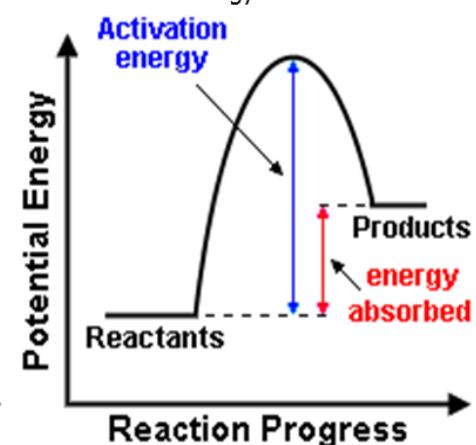
1 Conservation of energy	Energy is not created or destroyed , only transferred from one store to another
2 Exothermic	A reaction that transfers energy to the surroundings so the temperature of the surroundings increases , e.g. combustion and neutralisation reactions . Used in self-heating cans and hand warmers .
3 Endothermic	A reaction that takes in energy from the surroundings so the temperature of the surroundings decreases , e.g. thermal decomposition . Used in sports injury packs .
4 Activation energy	The energy needed for particles to successfully react .
5 Breaking bonds	Energy is needed to break bonds.
6 Forming bonds	Energy is released when bonds are formed.

7 Exothermic Energy Profile



Exothermic reaction

8 Endothermic Energy Profile



Endothermic reaction

9 Energy released from forming bonds is **greater than** the energy needed to break bonds. (HT)

10 Energy released from forming bonds is **less than** the energy needed to break bonds. (HT)

Weekly Biology Revision Plan 2020

Year 10

As a minimum you should be spending 30-45 minutes on Biology revision per week. The tasks listed below could take more than this if you do all of them so you will need to plan your time effectively. Regular revision is the key to success at GCSE, don't do too much in one go! We recommend approaching this revision as follows:

Primrose Kitten: Combined Science Biology Paper 1 <https://www.youtube.com/watch?v=mKYQ-K23Mr4>

GCSE AQA Combined Science TRILOGY <https://www.bbc.co.uk/bitesize/examspecs/z8r997h>

1. 10-15mins Read, cover and try to remember the information from You Tube and BBC Bitesize
2. 10-15mins creating your own revision resource to add to your folder (DO NOT COPY...TRANSFORM the information you have just read into something visual that you can remember)
3. 10-15mins Practicing application of what you have just revised, try exam questions and mark them using the mark schemes so you can correct your mistakes immediately!

Week beginning	Paper	Topic	Review (Read, cover, remember)	Revise and add to your revision folder	Teacher signed
Week 1	PAPER 1	Cell Biology	Cell structure	Create a set of flash cards for each cell and microscope type	
Week 2			Investigating cells		
Week 3			Cell Division	Create a mind map linking chromosomes, mitosis, stem cells and their uses	
Week 4		Organisation	Transport in and out of cells	Create a comparison table for diffusion, osmosis and active transport	
Week 5			Levels of organisation	Create a mnemonic to remember the order of the levels of organisation	
Week 6			Digestion	Draw an outline of the digestive system and label on what happens at each part (including which enzymes work at each part AND what they break down)	
Week 7		Blood and circulation	Blood and circulation	Make 3 posters to put up in your room - The make up of blood and the differences between the 3 different blood vessels - The heart and the direction of blood flow through it - Gas exchange in the lungs	
Week 8			Non-communicable diseases	Create flash cards for the different types of disease and their risk factors	
Week 9		Infection and response	Transport in plants	Draw a plant and a cross section of a leaf and label on the different types of transport. Include information of factors which may affect the speed of the transport.	
Week 10			Pathogens and disease	Create a table of diseases including information on pathogen, causes, symptoms and treatment.	
			Human defences against disease	Create a mind map of all the ways your body defends against	

		Treating diseases	pathogens and then link this to boosting immunity with vaccines and the different ways we can treat diseases.	
Week 11	Bioenergetics	Photosynthesis	Draw out the 3 rate of photosynthesis graphs and write a sentence to explain what is happening in each one.	
Week 12		Respiration and exercise	Make flashcards for each of the key subtitles on this page, be sure to include the 3 equations you need to learn.	

Weekly Physics Revision Plan 2020

Year 10

As a minimum you should be spending 30-45 minutes on Physics revision per week. The tasks listed below could take more than this if you do all of them so you will need to plan your time effectively. Regular revision is the key to success at GCSE, don't do too much in one go! We recommend approaching this revision as follows:

Primrose Kitten: Combined Science Physics Paper 1 <https://www.youtube.com/watch?v=xtw-Z0nllA4>

GCSE AQA Combined Science TRILOGY <https://www.bbc.co.uk/bitesize/topics/zqw77p3>

1. 10-15mins Read, cover and try to remember the information in the revision guide
2. 10-15mins creating your own revision resource to add to your folder (DO NOT COPY...TRANSFORM the information you have just read into something visual that can help you to remember)
3. 10-15mins Practicing application of what you have just revised, try exam questions and mark them using the mark schemes so you can correct your mistakes immediately!

Week beginning	Paper	Topic	Review (Read, cover, remember)	Revise and add to your revision folder	Teacher signed
Week 1	PAPER 1	Energy	Energy Stores and Transfers	Create a glossary of the key words in RED on this double page spread. Create flash cards containing the rearrangement triangles for the two energy equations you need to learn here.	
Week 2			Energy transfers and resources	Create sample flow diagrams for 3 or 4 different energy changes e.g. a car accelerating, a skateboarder at the top of a ramp, boiling water in a kettle. Make flashcards for all the different types of energy resources, include whether they are renewable or not and their advantages and disadvantages.	
Week 3		Electricity	Introduction to electricity	Create a poster of all the different components, their symbols and what they are used for. Include the charge equation, power equation, efficiency equation and energy transferred equation. Try and put the equations into the rearrangement triangles.	
Week 4			Circuits and resistance	Draw the 3 current/voltage graphs for resistors, filament lamps and diodes and annotate around them to describe what is happening.	
Week 5			Circuits and power	Draw a comparison table for series and parallel circuits. Include diagrams, information on current, resistance and potential difference.	
Week 6			Domestic uses of electricity	Draw and label/annotate diagrams for D.C, A.C (give examples of appliances that use both current types) and Wiring a three pin plug (explain what each part does).	
Week 7			Electrical energy in devices	Draw a flow chart to show how electricity gets from the power stations into our homes. Create a flash card with the advantages and disadvantages of overhead and underground cables.	
Week 8		Particle Model of Matter	Particle model of matter	Draw the particle model for solids, liquids and gasses and annotate the diagrams to include information on their properties. Draw a labelled diagram of the density practical and annotate with summarised method type bullet points.	
Week 9		Atomic Structure	Atoms and isotopes	Create a timeline for the development of the model of the atom. Include diagrams of previous models as well as the current accepted model and include the diagrams of the experiments that helped prove this new model.	
Week 10			Nuclear radiation	Create flash cards for each type of radiation to include their components (diagrams), hazards and what they may be absorbed by.	

			Create a glossary of key terms for the key words in RED and summarise radioactive contamination into a short paragraph.	
Week 11		Half life	<p>Draw the count rate graph for iodine-128 and explain what is half-life is and how you calculate it using the graph.</p> <p>Create a flash card for nuclear equations and then...PRACTICE, PRACTICE, PRACTICE!!</p>	

‘Seneca Learning’ Sign-Up Guide

Passcode: j5v9tvzq48

Step 1: Open an internet browser - *Any browser except Internet Explorer will work.*

Step 2: Go to [SenecaLearning.com](https://www.SenecaLearning.com)

Step 3: Click on “Get Started” or “Sign Up”

Step 4: Create your account - *If you don’t know your parent email, then type: N/A.*

Step 5: Click on “Classes & Assignments” - *You’ll find this in the top menu.*

Step 6: Click on “Join Class” - *It’s the green button in the top right corner.*

Step 7: Type the code from your teacher - *If you received a link instead, then open the link.*